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Chapter 15 Acids And Bases

Chapter 15: Acids and Bases
Acids and Bases.

Arrhenius Definitions:

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Acids And Bases Section 2: Arrhenius

acids- compounds that produce an increase in $[H^+]$ when dissolved in water. bases- compounds that produce an increase in $[OH^-]$ when dissolved in water Lewis Definitions: acids- electron pair acceptors.

Chapter 15: Acids and Bases

Chapter 15. Acids and Bases. GCh15-1.

Chapter 15. Acids and

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Acids And Bases

Section 2: Acids

- Bases. What we will learn:
- Bronsted acids and bases
 - Acid-base properties of water
 - pH - a measure of acidity
 - Strength of acids and bases
 - Weak acids (bases) and acid (base) ionization constant
 - Diprotic and polyprotic acids
 - Molecular structure and strength of acids
 - Acid-base properties of salts
 - Acid-base properties of oxides and hydroxides
 - Lewis

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acids and bases.

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Chapter 15. Acids and Bases

Chapter 15 Acids and Bases. STUDY. PLAY.

Chapter 15 main idea.

acids and bases

change the color of compound indicators.

... -Lewis acid-base

reaction is the

formation of one or

more covalent bonds

between an electron-

pair donor and an

electron pair acceptor,

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Section 2: Acids

although the 3 acid-base reactions differ, many compounds can be categorized as acids

...

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15 Drill: Identify the B-L acid and base in each of the following. Circle any amphoteric species acid acid base base Note: In both examples, water behaved as an acid or

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Section 2 Answers

a base. A species that can act as an acid or a base is called amphoteric. Bronsted-Lowry (B-L) Theory - Cont. acid base 1) $\text{HNO}_3 + \text{CO}_3^{2-} \Rightarrow \text{NO}_3^- + \text{HCO}_3^-$ 2) $\text{HPO}_4^{2-} + \text{H}_2\text{O} \Rightarrow \text{H} \dots$

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SECTION 15-1 O

OBJECTIVES List five

general properties of

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aqueous acids and bases. Name common binary acids and oxyacids, given their chemical formulas. List five acids commonly used in industry and the laboratory, and give two properties of each. Define acid and base according to Arrhenius's theory of ionization. Explain the differences

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Chapter 15 Aqueous Equilibria: Acids and Bases. Instructor's Resource Materials (Download only) for Chemistry, ...

Conjugate Acid-Base Pairs: Chemical species whose formulas differ only by one hydrogen ion, H^+ Brønsted-Lowry Acid: A substance that can transfer hydrogen ions, H^+ . In other words, a proton donor

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Chapter 15

Equilibria: Acids and Bases

acids = hydrochloric, nitric, sulfuric, phosphoric, ethanoic (acetic), carbonic.

bases = potassium hydroxide, sodium hydroxide, calcium hydroxide, magnesium hydroxide. Bronsted-Lowry acids and bases.

acid - molecule or ion that is a proton donor (a proton is a hydrogen ion, H^+) base -

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molecule or ion that is a proton acceptor.

Chapter 15 (Acids and Bases)

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Acids and Bases Book

Notes. Sec 15.1:

Heartburn: Sec 15.2:

The Nature of Acids

and Bases: Properties

of Acids-sour taste,

ability to dissolve many

metals, ability to turn

blue litmus paper red,

and ability to neutralize

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bases
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help you improve your
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Chapter 15 : Acids
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and Bases
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Chapter 15 Acids and Bases. strong acid. strong base. weak acid. weak base. an acid that ionizes completely in solvent. a base that ionizes completely in a solvent. an acid that releases few hydrogen ions in aqueous solution. a base that releases few hydroxide ions in aqueous solution.

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15 Flashcards and**

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Chapter 15 - Acids,
Bases, & Salts. Lewis
Acid. Lewis Base.

Bronsted Lowry Acid.

Bronsted Lowry Base.

Anything that accepts
a pair of electrons.

Anything that donates
a pair of electrons in a
reaction. Anything that
donates a proton, H^+ ,
in a reaction. Anything
that accepts a proton
in a reaction.

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acids bases salts
chapter 15
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Chemistry: A Molecular Approach, 3e (Tro).

Chapter 15 Acids and Bases Multiple Choice

Questions 1) _____ is found in carbonated beverages due to the reaction of carbon dioxide with water.

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An acid is a substance
that, when dissolved in

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Section 2 Answers

water, increases the concentration of hydrogen ions. A base is a substance that, when dissolved in water, increases the concentration of hydroxide ions.

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- An acid is a substance that, when dissolved in water, increases the concentration of

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hydrogen ions. - A base is a substance that, when dissolved in water, increases the concentration of hydroxide ions.

Chapter 15 Acids and Bases

Acids and Bases:

Chapter 14 & 15 . HW:

- Read Ch 14: •Fill in as much of the acid base table as you can, as you read . Acid base conductivity and reactivity Conductivity,

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Acids And Bases

Reac'tivity,,

Hydrochloric*acid* high

high Acidic*acid* **

low* medium

Dis4lled*water* none*

none*

Ammoniumhydroxide

med* none*

Sodiumhydroxide* high

none* *

Acids and Bases:

Chapter 14 & 15

Strong acids and bases are often powerful, dangerous substances.

(H₂SO₄) will

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decompose sugar into black charcoal. Nitric acid (HNO_3) reacts with many metals, and hydrochloric acid (HCl) will eat...

o. chapter 15 acids and bases UPDATED by ...

CHM 112 Chapter 15
Worksheet: Acids and Bases Name: _____ The appropriate K_a or K_b values can be found in your textbook. ...

Hydroxylamine is a

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Solution 2 Answers

weak base. A 0.15 M solution of hydroxylamine has a pH of 10.11. What is the K_b for this base? $K_b \dots$ Lewis acid Lewis Base \leftrightarrow Lewis Acid-Base Product I

CHM 112 Chapter 15 Worksheet: Acids and Bases Name: The ...

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Problems - Page 659
15.78 including work
step by step written by
community members
like you. Textbook
Authors: Ebbing,
Darrell; Gammon,
Steven D., ISBN-10:
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